Aim:  
To see how temp affects the rate of the reaction between HCl and CaCO3

Method:  
Your set up (diagram & what you did)

Conclusion:  
The higher the temp the faster the reaction between HCl and CaCO3

Discussion:  
For a reaction to occur particles must collide with sufficient energy and the correct orientation. In this case s we heated the HCl the particles move faster and have more energy. Since they are moving faster they will collide more often. Also the collisions will have more energy. Therefore there will be more successful collisions per minute. So the reaction rate will be faster at higher temperature.

Evaluation:  
To make it a fair test we:

* Used the same amount of acid (1o mL) because if we had more acid it might have gone faster
* Used same sized chips because different sizes would have a different surface area and that might affect the rate
* Used the same concentration of HCl each time because a higher conc is likely to make the reaction go faster
* Used the same equipment each time

To make it reliable:

we carried out 6 repeats so that any variation in results could be smoothed out when we looked at the average.